



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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CANDIDATE  
NUMBER

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**CHEMISTRY**

**0620/21**

Paper 2

**October/November 2013**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may need to use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 16.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

This document consists of **14** printed pages and **2** blank pages.



- 1 (a) Choose from the list of elements below to answer the following questions.

For  
Examiner's  
Use

calcium  
helium  
iodine  
nickel  
nitrogen  
sodium  
sulfur

Each element can be used once, more than once or not at all.

Which element:

- (i) is an element present in most fertilisers, ..... [1]  
 (ii) is in Group VI of the Periodic Table, ..... [1]  
 (iii) is in Period 5 of the Periodic Table, ..... [1]  
 (iv) has a single electron shell containing two electrons, ..... [1]  
 (v) is a transition element, ..... [1]  
 (vi) forms ions with a single negative charge? ..... [1]

- (b) What is the meaning of the term *element*?

.....  
 ..... [1]

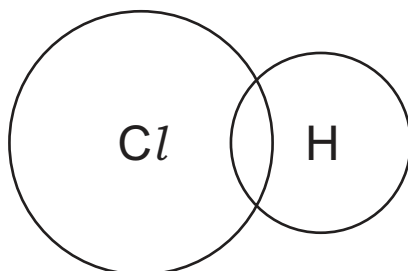
- (c) Many of the elements in the Periodic Table have metallic properties.  
 Describe **three** physical properties which are typical of most metals.

1. ....  
 2. ....  
 3. .... [3]

[Total: 10]

2 Hydrogen chloride is an acidic gas.

- (a) (i) Complete the dot and cross diagram to show the electronic structure of hydrogen chloride.



[2]

- (ii) Is hydrogen chloride a covalent or an ionic compound?  
Give a reason for your answer.

..... [1]

- (b) Hydrogen chloride reacts with water to form hydrochloric acid.  
Which one of the following is the most likely pH of hydrochloric acid?  
Put a ring around the correct answer.

pH2      pH7      pH9      pH14

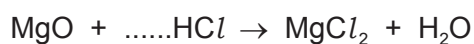
[1]

- (c) Hydrochloric acid reacts with both metal oxides and carbonates.

- (i) Complete the word equation for the reaction of hydrochloric acid with calcium carbonate.

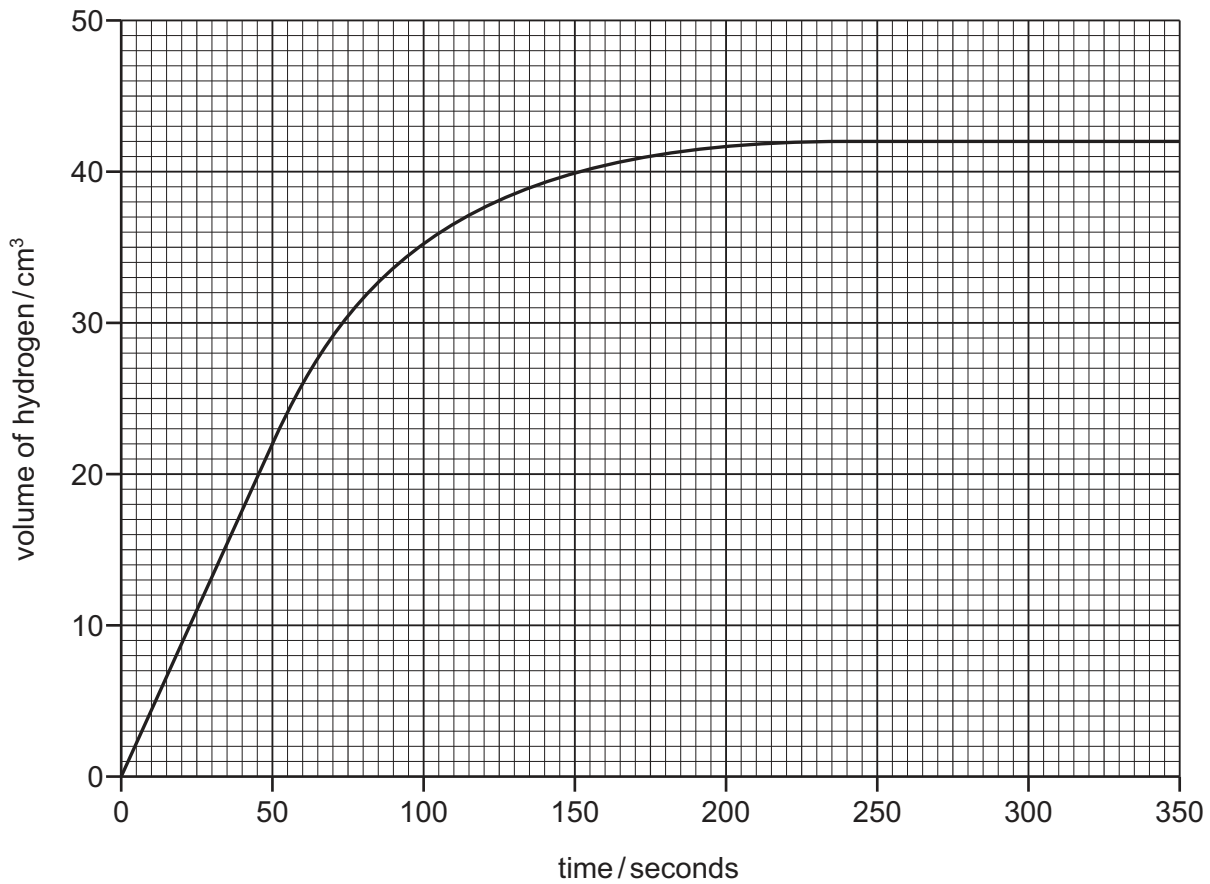
hydrochloric + calcium → ..... + ..... + .....  
acid            carbonate  
.....            ..... [3]

- (ii) Complete the symbol equation for the reaction of magnesium oxide with hydrochloric acid. Name the salt which is formed.



name of salt ..... [2]

- (d) A student reacted magnesium with hydrochloric acid to find out how concentration affects the rate of reaction. The magnesium was in excess. He measured the volume of hydrogen produced at various time intervals. The graph shows his results.



- (i) At what time had the reaction just finished?

..... [1]

- (ii) What volume of hydrogen gas is given off during the first 50 seconds of the reaction?

volume of hydrogen ..... cm<sup>3</sup> [1]

- (iii) The student repeated the experiment.

State **two** factors, apart from the concentration of hydrochloric acid, that should be kept constant when repeating the experiment.

1. ....

2. .... [2]

[Total: 13]

3 Organic compounds can be put into groups called homologous series.

(a) Complete the following sentences about organic compounds and homologous series. Use words from the list below.

**carbon**      **chlorine**      **different**      **elements**      **functional**  
**hydrocarbon**      **hydrogen**      **oxide**      **similar**      **sulfur**

Organic compounds usually contain atoms of ..... and .....

Each homologous series contains compounds with ..... chemical properties due to the presence of the same ..... group. [4]

(b) Ethanol belongs to the alcohol homologous series.

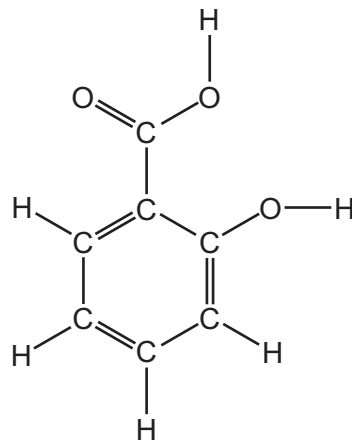
(i) Draw the structure of ethanol, showing all atoms and bonds.

[2]

(ii) State the name of the **two** compounds formed when ethanol burns in excess air.

..... and ..... [2]

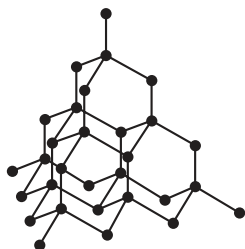
- (c) Salicylic acid is used to make aspirin.  
The structure of salicylic acid is shown below.



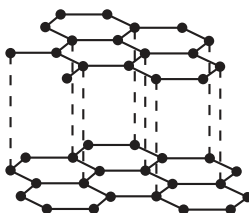
- (i) On this structure, put a ring around the carboxylic acid functional group. [1]
- (ii) How many carbon atoms are there in one molecule of salicylic acid?  
..... [1]
- (iii) When making drugs and medicines, it is important that the chemicals used are pure.  
State **one** other area of everyday life where purity is important.  
..... [1]

[Total: 11]

4 The structures of diamond and graphite are shown below.



diamond



graphite

• = carbon atom

(a) Describe the similarities and differences between these structures.

.....

.....

.....

.....

.....

..... [4]

(b) Graphite burns in excess air to form carbon dioxide.  
Describe a test for carbon dioxide.

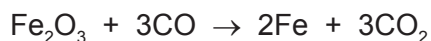
test .....

result ..... [2]

(c) When graphite is burnt in a limited supply of air, carbon monoxide is formed.  
State **one** adverse effect of carbon monoxide on health.

..... [1]

(d) In the blast furnace for the production of iron, carbon monoxide reduces iron(III) oxide.



How does this equation show that carbon monoxide is acting as a reducing agent?

..... [1]

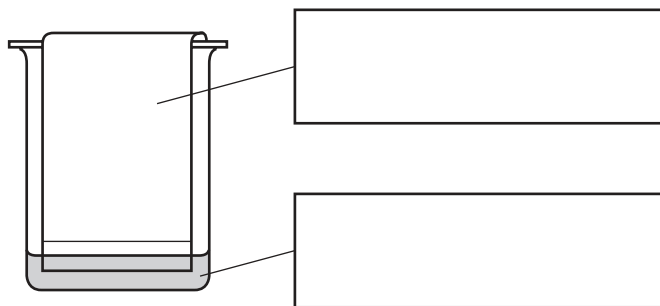
(e) Iron(III) oxide and coke (carbon) are raw materials used in the production of iron.  
State the names of **two** other raw materials used in the blast furnace for the production of iron.

1. ....

2. .... [2]

[Total: 10]

- 5 Many plants contain coloured pigments.  
A student crushes some plant leaves in alcohol to extract the pigments.  
She then separates the pigments using the apparatus shown below.



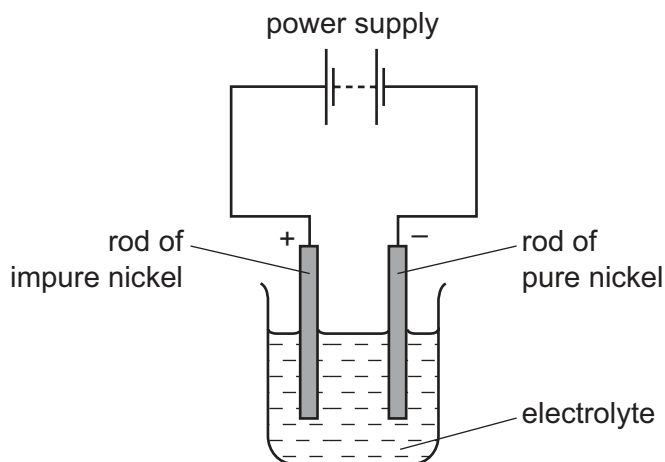
(a) Write the correct labels in the boxes in the diagram above. [2]

(b) Draw an **X** on the diagram above to show where a drop of the pigment solution is placed at the start of the experiment. [1]

(c) After leaving the apparatus for half an hour, the pigments separated from each other. State the name given to this method of separating pigments.

..... [1]

(d) Some plants can absorb nickel from the ground. The nickel can then be extracted from the plants and purified by electrolysis.



(i) Which one of the following is the most suitable electrolyte for this electrolysis. Tick **one** box.

- aqueous copper(II) sulfate
- aqueous nickel(II) sulfate
- solid nickel(II) sulfate
- water

[1]



- (ii) Which one of the following elements is most likely to be formed at the negative electrode during this electrolysis?  
Put a ring around the correct answer.

**chlorine                  nickel                  sulfur                  oxygen**

[1]

- (iii) The positive electrode is called the anode.  
State the name of the negative electrode.

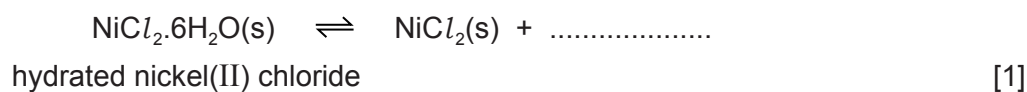
..... [1]

- (e) Electroplating is used to put a thin layer of one metal on top of another by electrolysis.  
Give **two** reasons for electroplating metals.

1. ....

2. .... [2]

- (f) (i) Hydrated nickel(II) chloride is green in colour.  
When hydrated nickel(II) chloride is heated gently, it changes colour from green to white.  
Complete the symbol equation for this reaction.



- (ii) What does the sign  $\rightleftharpoons$  mean?

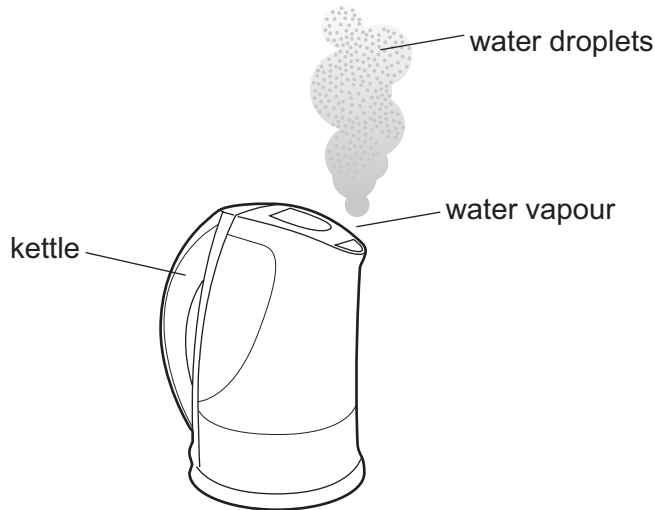
..... [1]

- (iii) How can you obtain a sample of green nickel(II) chloride starting with white nickel(II) chloride?

..... [1]

[Total: 12]

6 The diagram shows a kettle of boiling water.



As the water vapour cools it turns back to water droplets.

(a) Describe this change of state in terms of the kinetic particle theory.  
In your answer, include

- the difference in the closeness of the water molecules as the water vapour changes to water,
- the difference in the motion of the water molecules as the water vapour changes to water.

.....  
.....  
.....  
.....  
..... [4]

(b) Water is a common solvent in the laboratory.

(i) What is meant by the term *solvent*?

..... [1]

(ii) State the name of the solvent whose formula is  $C_2H_5OH$ .

..... [1]

(c) When ammonium chloride dissolves in water the temperature of the solution falls.  
State the name of the energy change which results in the temperature falling.

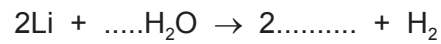
..... [1]

- (d) Which one of the following conducts electricity.  
Tick **one** box.

aqueous ammonium chloride	<input type="checkbox"/>
solid ammonium chloride	<input type="checkbox"/>
ammonia gas	<input type="checkbox"/>
chlorine gas	<input type="checkbox"/>

[1]

- (e) (i) Complete the symbol equation for the reaction of lithium with water to form lithium hydroxide and hydrogen.



[2]

- (ii) When 14 g of lithium react with water, 4 g of hydrogen are formed.  
Calculate the mass of hydrogen formed when 70 g of lithium react with water.

[1]

[Total: 11]

7 The table shows some properties of seven different substances.

substance	density /g per cm <sup>3</sup>	relative strength	relative electrical conductivity	relative thermal conductivity
aluminium	2.7	15	42	200.0
ceramic	2.5	15	does not conduct	1.6
copper	8.9	20	63	385.0
iron	7.9	25	11	80.0
lead	11.4	15	5	38.0
poly(ethene)	0.9	1	does not conduct	0.3
steel	7.8	90	2	25.0

(a) Use the information in this table to answer the following questions.

(i) Which substance is the best conductor of heat?

..... [1]

(ii) Suggest why copper is preferred to iron for electrical wiring in houses.

..... [1]

(iii) What property of ceramic makes it a good electrical insulator?

..... [1]

(iv) Which pure metal in the table conducts electricity least well?

..... [1]

(v) Suggest why steel rather than iron is used in making machinery.

..... [1]

(vi) Which metal in the table is the most dense?

..... [1]

(b) A solution of a metal salt reacts with aqueous sodium hydroxide to form a white precipitate. The white precipitate is soluble in excess aqueous sodium hydroxide.

(i) Which one of the following ions is most likely to be present in the salt?  
Put a ring around the correct answer.

**calcium**      **copper(II)**      **iron(II)**      **zinc**

[1]

(ii) State the name of the white precipitate.

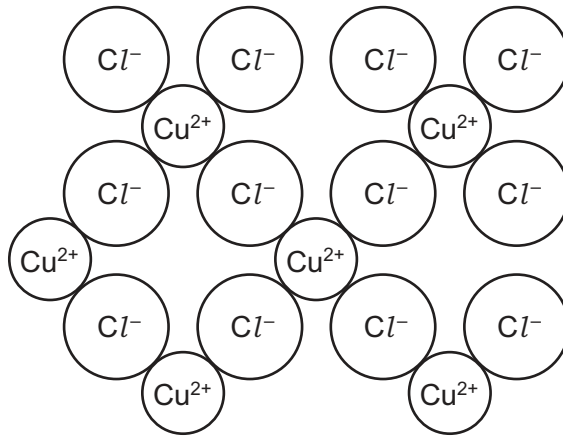
..... [1]

- (c) Copper(II) chloride can be made by the action of hydrochloric acid on copper(II) oxide. Put the statements, **A**, **B**, **C** and **D**, about this preparation in the correct order.

- A** Leave the saturated solution to crystallise.  
**B** Filter the solution to remove excess copper(II) oxide.  
**C** Add excess copper(II) oxide to hydrochloric acid and warm.  
**D** Evaporate the filtrate to the crystallisation point.

..... [1]

- (d) The structure of copper(II) chloride is shown below.



Write the simplest formula for copper(II) chloride.

..... [1]

- (e) Suggest the product formed at each electrode when molten copper(II) chloride is electrolysed.

at the positive electrode .....

at the negative electrode ..... [2]

- (f) When copper(II) chloride is heated strongly, a gas is given off. The gas is green in colour and bleaches litmus paper. State the name of this gas.

..... [1]

[Total: 13]





**DATA SHEET**  
**The Periodic Table of the Elements**

		Group																																																																																														
I	II	III	IV	V	VI	VII	0																																																																																									
7 <b>Li</b> Lithium 3	9 <b>Be</b> Beryllium 4	1 <b>H</b> Hydrogen 1	11 <b>B</b> Boron 5	12 <b>C</b> Carbon 6	13 <b>Al</b> Aluminium 13	14 <b>N</b> Nitrogen 7	15 <b>O</b> Oxygen 8	16 <b>F</b> Fluorine 9	17 <b>Ne</b> Neon 10	18 <b>Ar</b> Argon 18	19 <b>K</b> Potassium 19	20 <b>Ca</b> Calcium 20	21 <b>Sc</b> Scandium 21	22 <b>Ti</b> Titanium 22	23 <b>V</b> Vanadium 23	24 <b>Cr</b> Chromium 24	25 <b>Mn</b> Manganese 25	26 <b>Fe</b> Iron 26	27 <b>Co</b> Cobalt 27	28 <b>Ni</b> Nickel 28	29 <b>Cu</b> Copper 29	30 <b>Zn</b> Zinc 30	31 <b>Ga</b> Gallium 31	32 <b>Ge</b> Germanium 32	33 <b>As</b> Arsenic 33	34 <b>Se</b> Selenium 34	35 <b>Br</b> Bromine 35	36 <b>Kr</b> Krypton 36	37 <b>Rb</b> Rubidium 37	38 <b>Sr</b> Strontium 38	39 <b>Y</b> Yttrium 39	40 <b>Zr</b> Zirconium 40	41 <b>Nb</b> Niobium 41	42 <b>Mo</b> Molybdenum 42	43 <b>Tc</b> Technetium 43	44 <b>Ru</b> Ruthenium 44	45 <b>Rh</b> Rhodium 45	46 <b>Pd</b> Palladium 46	47 <b>Ag</b> Silver 47	48 <b>Cd</b> Cadmium 48	49 <b>In</b> Indium 49	50 <b>Sn</b> Tin 50	51 <b>Sb</b> Antimony 51	52 <b>Te</b> Tellurium 52	53 <b>I</b> Iodine 53	54 <b>Xe</b> Xenon 54	55 <b>Cs</b> Caesium 55	56 <b>Ba</b> Barium 56	57 <b>La</b> Lanthanum 57	72 <b>Hf</b> Hafnium 72	73 <b>Ta</b> Tantalum 73	74 <b>W</b> Tungsten 74	75 <b>Re</b> Rhenium 75	76 <b>Os</b> Osmium 76	77 <b>Ir</b> Iridium 77	78 <b>Pt</b> Platinum 78	79 <b>Au</b> Gold 79	80 <b>Hg</b> Mercury 80	81 <b>Tl</b> Thallium 81	82 <b>Pb</b> Lead 82	83 <b>Bi</b> Bismuth 83	84 <b>Po</b> Polonium 84	85 <b>At</b> Astatine 85	86 <b>Rn</b> Radon 86	87 <b>Fr</b> Francium 87	88 <b>Ra</b> Radium 88	89 <b>Ac</b> Actinium 89	90 <b>Th</b> Thorium 90	91 <b>Pa</b> Protactinium 91	92 <b>U</b> Uranium 92	93 <b>Np</b> Neptunium 93	94 <b>Pu</b> Plutonium 94	95 <b>Am</b> Americium 95	96 <b>Cm</b> Curium 96	97 <b>Bk</b> Berkelium 97	98 <b>Cf</b> Californium 98	99 <b>Es</b> Einsteinium 99	100 <b>Fm</b> Fermium 100	101 <b>Md</b> Mendelevium 101	102 <b>No</b> Nobelium 102	103 <b>Lr</b> Lawrencium 103	104 <b>Rf</b> Rutherfordium 104	105 <b>Db</b> Dubnium 105	106 <b>Sg</b> Seaborgium 106	107 <b>Bh</b> Bohrium 107	108 <b>Hs</b> Hassium 108	109 <b>Mt</b> Meitnerium 109	110 <b>Ds</b> Darmstadtium 110	111 <b>Rg</b> Roentgenium 111	112 <b>Cn</b> Copernicium 112	113 <b>Nh</b> Nihonium 113	114 <b>Fl</b> Flerovium 114	115 <b>Mc</b> Moscovium 115	116 <b>Lv</b> Livermorium 116	117 <b>Ts</b> Tennessine 117	118 <b>Og</b> Oganesson 118
												140 <b>Ce</b> Cerium 58	141 <b>Pr</b> Praseodymium 59	142 <b>Nd</b> Neodymium 60	143 <b>Pm</b> Promethium 61	144 <b>Nd</b> Neodymium 60	145 <b>Sm</b> Samarium 62	146 <b>Eu</b> Europium 63	147 <b>Gd</b> Gadolinium 64	148 <b>Tb</b> Terbium 65	149 <b>Dy</b> Dysprosium 66	150 <b>Ho</b> Holmium 67	151 <b>Er</b> Erbium 68	152 <b>Tm</b> Thulium 69	153 <b>Yb</b> Ytterbium 70	154 <b>Lu</b> Lutetium 71	155 <b>Yt</b> Ytterbium 70	156 <b>Hf</b> Hafnium 72	157 <b>Ta</b> Tantalum 73	158 <b>W</b> Tungsten 74	159 <b>Re</b> Rhenium 75	160 <b>Os</b> Osmium 76	161 <b>Pt</b> Platinum 78	162 <b>Au</b> Gold 79	163 <b>Hg</b> Mercury 80	164 <b>Tl</b> Thallium 81	165 <b>Pb</b> Lead 82	166 <b>Bi</b> Bismuth 83	167 <b>Po</b> Polonium 84	168 <b>At</b> Astatine 85	169 <b>Rn</b> Radon 86	170 <b>Fr</b> Francium 87	171 <b>Ra</b> Radium 88	172 <b>Ac</b> Actinium 89	173 <b>Th</b> Thorium 90	174 <b>Pa</b> Protactinium 91	175 <b>U</b> Uranium 92	176 <b>Np</b> Neptunium 93	177 <b>Pu</b> Plutonium 94	178 <b>Am</b> Americium 95	179 <b>Cm</b> Curium 96	180 <b>Bk</b> Berkelium 97	181 <b>Cf</b> Californium 98	182 <b>Es</b> Einsteinium 99	183 <b>Fm</b> Fermium 100	184 <b>Md</b> Mendelevium 101	185 <b>No</b> Nobelium 102	186 <b>Lr</b> Lawrencium 103																																						

\*58-71 Lanthanoid series  
†90-103 Actinoid series

**Key**  

a	<b>X</b>
b	

 a = relative atomic mass  
 x = atomic symbol  
 b = proton (atomic) number

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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